

Critical Item 8 – General Chemistry II

copy 1

NAME _____

What is the Cu^+ concentration if a solution is $6.7 \times 10^{-3} \text{ M}$ in S^{2-} in contact with precipitated Cu_2S .
 $K_{sp} = 2.0 \times 10^{-47}$. Give the right number to 2 significant figures.

Write the equilibrium reaction here: _____

Show your calculation here clearly

 $[\text{Cu}^+] =$ _____**Critical Item 8 – General Chemistry II**

copy 2

NAME _____

What is the Fe^{2+} concentration if a solution is $1.1 \times 10^{-3} \text{ M}$ in OH^- in contact with precipitated $\text{Fe}(\text{OH})_2$.
 $K_{sp} = 9.0 \times 10^{-35}$. Give the right number to 2 significant figures.

Write the equilibrium reaction here: _____

Show your calculation here clearly

 $[\text{Fe}^{2+}] =$ _____**Critical Item 8 – General Chemistry II**

copy 3

NAME _____

What is the Ba^{2+} concentration if a solution is $1.9 \times 10^{-2} \text{ M}$ in F^- in contact with precipitated BaF_2 .
 $K_{sp} = 1.7 \times 10^{-6}$. Give the right number to 2 significant figures.

Write the equilibrium reaction here: _____

Show your calculation here clearly

 $[\text{Ba}^{2+}] =$ _____**KEY**copy 1 $[\text{Cu}_2\text{S}] = 5.5 \times 10^{-23}$ copy 2 $[\text{Fe}(\text{OH})_2] = 7.3 \times 10^{-29}$ copy 3 $[\text{BaF}_2] = 4.7 \times 10^{-3}$