

$$F = 96\,487 \text{ C mol}^{-1}$$

How many grams of copper will be plated after 7.73 hours if a constant current of 14.7 amps is used? The reaction is: $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$

$$\text{ANS: } \underline{\hspace{1cm}} 134.7 \underline{\hspace{1cm}} \text{ g}$$

How many grams of chromium will be plated after 3.66 hours if a constant current of 17.8 amps is used? The reaction is: $\text{Cr}^{3+} + 3\text{e}^- \rightarrow \text{Cr}$

$$\text{ANS: } \underline{\hspace{1cm}} 42.1 \underline{\hspace{1cm}} \text{ g}$$

How many grams of chromium will be plated after 7.30 hours if a constant current of 12.9 amps is used? The reaction is: $\text{Cr}^{3+} + 3\text{e}^- \rightarrow \text{Cr}$

$$\text{ANS: } \underline{\hspace{1cm}} 60.9 \underline{\hspace{1cm}} \text{ g}$$

How many grams of nickel will be plated after 7.73 hours if a constant current of 16.6 amps is used? The reaction is: $\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$

$$\text{ANS: } \underline{\hspace{1cm}} 140.5 \underline{\hspace{1cm}} \text{ g}$$

How many grams of nickel will be plated after 7.04 hours if a constant current of 10.7 amps is used? The reaction is: $\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$

$$\text{ANS: } \underline{\hspace{1cm}} 82.5 \underline{\hspace{1cm}} \text{ g}$$

How long would it take to plate out 599 grams of nickel if a constant current of 11.1 amps is used? The reaction is: $\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$

$$\text{ANS: } \underline{\hspace{1cm}} 49.3 \underline{\hspace{1cm}} \text{ hours}$$

How long would it take to plate out 848 grams of chromium if a constant current of 10.2 amps is used? The reaction is: $\text{Cr}^{3+} + 3\text{e}^- \rightarrow \text{Cr}$

$$\text{ANS: } \underline{\hspace{1cm}} 128.6 \underline{\hspace{1cm}} \text{ hours}$$

How long would it take to plate out 200 grams of copper if a constant current of 14.7 amps is used? The reaction is: $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$

$$\text{ANS: } \underline{\hspace{1cm}} 11.5 \underline{\hspace{1cm}} \text{ hours}$$

What current is required to plate 57.5 grams of copper if the plating time is 11.9 hours? The reaction is: $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$

$$\text{ANS: } \underline{\hspace{1cm}} 4.08 \underline{\hspace{1cm}} \text{ amps}$$

What current is required to plate 83.1 grams of chromium if the plating time is 11.2 hours? The reaction is: $\text{Cr}^{3+} + 3\text{e}^- \rightarrow \text{Cr}$

$$\text{ANS: } \underline{\hspace{1cm}} 11.47 \underline{\hspace{1cm}} \text{ amps}$$