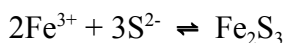
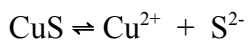


Calculate the equilibrium constant at 298 K for the following reaction:



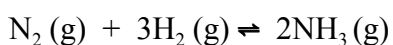
$$\text{Ans} = 1.04 \times 10^{39}$$

Calculate the equilibrium constant at 298 K for the following reaction:



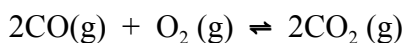
$$\text{Ans} = 3.1 \times 10^{-26}$$

Calculate the equilibrium constant at 298 K for the following reaction:



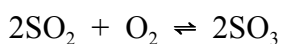
$$\text{Ans} = 6.1 \times 10^5$$

Calculate the equilibrium constant at 298 K for the following reaction:



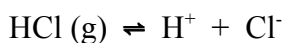
$$\text{Ans} = 1.5 \times 10^{90}$$

Calculate the equilibrium constant at 298 K for the following reaction:



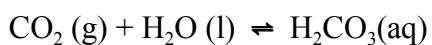
$$\text{Ans} = 6.94 \times 10^{24}$$

Calculate the equilibrium constant at 298 K for the following reaction:



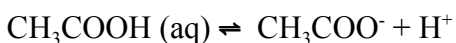
$$\text{Ans} = 2.1 \times 10^6$$

Calculate the equilibrium constant at 298 K for the following reaction:



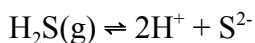
$$\text{Ans} = 3.52 \times 10^{-2}$$

Calculate the equilibrium constant at 298 K for the following reaction:



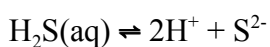
$$\text{Ans} = 1.74 \times 10^{-5}$$

Calculate the equilibrium constant at 298 K for the following reaction:



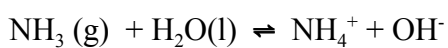
$$\text{Ans} = 1.2 \times 10^{-21}$$

Calculate the equilibrium constant at 298 K for the following reaction:



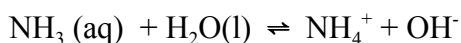
$$\text{Ans} = 1.4 \times 10^{-20}$$

Calculate the equilibrium constant at 298 K for the following reaction:



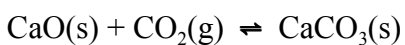
$$\text{Ans} = 1.06 \times 10^{-3}$$

Calculate the equilibrium constant at 298 K for the following reaction:



$$\text{Ans} = 1.84 \times 10^{-5}$$

Calculate the equilibrium constant at 298 K for the following reaction:



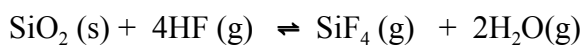
$$\text{Ans} = 7.4 \times 10^{22}$$

Calculate the equilibrium constant at 298 K for the following reaction:



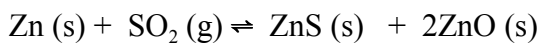
$$\text{Ans} = 3.4 \times 10^6$$

Calculate the equilibrium constant at 298 K for the following reaction:



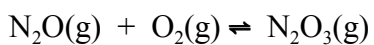
$$\text{Ans} = 1.19 \times 10^{14}$$

Calculate the equilibrium constant at 298 K for the following reaction:



$$\text{Ans} = 4.2 \times 10^{98}$$

Calculate the equilibrium constant at 298 K for the following reaction:



$$\text{Ans} = 4.4 \times 10^{-8}$$